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**TERMS TO BE DEFINED & KEPT IN MIND (Reflection)****PHASE EQUILIBRIUM**

1. Phase(P)
2. Component(C)
3. Degrees of freedom(F)
4. Gibb's phase rule
5. Clapeyron equation
6. Clausius-Clapeyron equation
7. Phase equilibria
8. Phase transitions
9. Melting, Freezing, Vaporization, Boiling , Sublimation
10. Melting point , Freezing point , Boiling point
11. Triple point
12. Invariant point
13. Meta stable equilibrium
14. Condensed System
15. Reduced phase rule
16. Eutectic temperature
17. Eutectic composition
18. Desilverisation of Lead
19. Congruent melting
20. Incongruent melting.
21. Congruent melting point
22. Incongruent melting point.
23. Peritectic Change
24. Metatectic Change
25. Freezing mixture
26. Partially miscible pair
27. Binodal curve
28. Tie line

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**SOLUTIONS & COLLIGATIVE PROPERTIES**

29. Solvent, Solute , Solution
30. Ideal solutions
31. Vapor pressure
32. Raoult's law
33. Positive deviation from ideal behavior
34. Negative deviation from ideal behavior
35. Distillation
36. Fractional distillation
37. Azeotropic boiling point
38. Azeotropic composition
39. Miscibility temperature
40. Critical solution temperature (CST)
41. Upper Critical solution temperature (UCST)
42. Lower Critical solution temperature (LCST)
43. Steam distillation
44. Henry's law
45. Lowering of vapor pressure
46. Relative Lowering of vapor pressure
47. Mole fraction ( $x$ )
48. Molality ( $m$ )
49. Molarity ( $M$ )
50. Elevation of boiling point
51. Ebullioscopic constant
52. Depression of freezing point
53. Cryoscopic constant
54. Osmotic pressure
55. Van't Hoff's factor
56. Nernst Distribution law
57. Distribution (Partition) coefficient
58. Efficiency of extraction.